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1. Gram-Mole Conversions. 2. Solving Stoichiometry Problems ... The mole tunnel is referring to a chemical equation like the one below. For this example, if you want to make aluminum hydroxide, you need one unit of aluminum chloride for every three units of sodium hydroxide. You must convert it to mass ...

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PRACTICE PACKET: (Unit 6 Moles & Stoichiometry) (7 www.mrpalermo.com) Use the formula below to answer questions! 8?13  
 $! Fe_2O_3 + 3CO \rightarrow 2Fe + 2CO_2$  (8.  
If 3.00 moles of Iron ...

## Practice Packet (Unit 6: Moles & Stoichiometry

$0.012 \text{ mol CO}_2 \cdot \frac{1}{3} = 0.0040 \text{ mol Fe}$   
 $0.012 \text{ moles CO}_2 \cdot 44 \text{ g/mol} = 0.53 \text{ g CO}_2$   
 $0.0052 - 0.0040 = 0.0012 \text{ mol left CO}_2$   
 $0.0012 \text{ mol} \cdot 192 \text{ g/mol} = 0.23 \text{ g left.}$

Stoichiometry Theoretical Yield • The theoretical yield is the amount of product that can be made - In other words it's the amount of product possible from stoichiometry. The "perfect ...

## Chapter 3 Stoichiometry - Chemistry

How many moles of water are required to react with 1.25 moles of iron

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according to the following reaction?  $3\text{Fe} + 4\text{H}_2\text{O} \rightarrow \text{Fe}_3\text{O}_4 + 4\text{H}_2$  View Answer Aluminum is oxidized in air to become aluminum ...

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