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CHAPTER 14 REVIEW Acids and Bases
SECTION 2 SHORT ANSWER Answer the
following questions in the space
provided. 1. a. Write the two equations
that show the two-stage ionization of
sulfurous acid in water. stage 1: $\text{H}_2\text{SO}_3(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{H}_3\text{O}^+(\text{aq}) + \text{HSO}_3^-(\text{aq})$

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stage 2: $\text{HSO}_3^- (\text{aq}) + \text{H}_2\text{O} (\text{l}) \rightleftharpoons \text{H}_2\text{SO}_3 (\text{aq}) + \text{OH}^- (\text{aq})$ b.

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SHORT ANSWER Answer the following
questions in the space provided. 1. a.
Write the two equations that show the
two-stage ionization of sulfurous acid in
water. stage 1: $\text{H}_2\text{SO}_3(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons$
 $\text{H}_3\text{O}^+(\text{aq}) + \text{HSO}_3^-(\text{aq})$ b. Which stage of

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ionization usually produces more ions?
Explain your answer.

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Know the meanings of and be able to
apply the following terms: Bronsted-
Lowry acid strong acid concentrated
chemical ...

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Chapter 14 Acids and Bases. STUDY.
PLAY. Properties of acids. taste sour,
contain H and some react with active
metals to liberate H_2 gas, react with
bases to produce salts and water, are

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electrolytes, change the color of dyes known as "acid-base indicators" (litmus paper, phenolphthalein) ... Chapter 20 Review Sheet. 23 terms. Chem Acid/Base ...

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Acids and Bases: Chapter 14 & 15 . HW:

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- Read Ch 14: • Fill in as much of the acid base table as you can, as you read . Acid base conductivity and reactivity
- Conducvity, Reac'vity,, ... 3. Work on ch 14-15 review (due Wednesday, test Friday!) • Solutions in which [H

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Chapter 14 - Acids and Bases.

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11th Grade. Created. ... Definition. 1) sour taste, corrosive 2) characteristically changes acid-base indicator 3) react with active metals to form hydrogen gas 4) react with bases to form salt and water 5) conduct electric current: Term.

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Acid is H^+ DONOR, Base is H^+ ACCEPTOR. In the reaction, the acid molecule gives an H^+ to the base molecule.

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CHAPTER 14 ACIDS AND BASES 3 Weak acids are only partially dissociated in

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water. We say that the equilibrium lies far to the left, thus giving values for $K_a < 1$. (We have mostly reactants at equilibrium and few products present).

CHAPTER FOURTEEN ACIDS AND BASES

Chapter 14: Acids and Bases Ch 14 Page 1
Arrhenius Base -a hydroxide

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containing compound that dissociates to produce OH⁻ when dissolved in water
 $\text{NaOH(aq)} \rightarrow \text{Na}^+(\text{aq}) + \text{OH}^-(\text{aq})$ Bases that contain OH⁻ are ionic compounds. Ionic substances dissociate in water.

Chapter 14: Acids and Bases

CHAPTER 14 REVIEW. Acids and Bases.
SHORT ANSWER Answer the following

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questions in the space provided. 1.
Answer the following questions
according to the Brønsted-Lowry
definitions of acids and bases: HSO_3^- . 3 -.
a. What is the conjugate base of H.

CHAPTER 14 REVIEW Acids and Bases - Weebly

Chapter Fourteen: Acids and Bases

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Chapter interactives: Acids/bases and indicators
Chemistry battleship
Chemistry jokes
pH simulation
Chapter Homework: Section 1: Chapter review 1 thru 7.
Section 2: Chapter review 12, 13.
Section 3: Chapter review 19 thru 25.

Chapter Fourteen [Acids and Bases] - Wattsburg

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Modern Chemistry 121 Acids and Bases
CHAPTER 14 REVIEW Acids and Bases
SECTION 3 SHORT ANSWER Answer the
following questions in the space
provided. 1. Answer the following
questions according to the Brønsted-
Lowry definitions of acids and bases:

_____ a. What is the conjugate base of H
2SO 3? _____ b.

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CHAPTER 14 REVIEW Acids and Bases

has acids and conjugate bases in order of strength like Table 14.1 in McMurry Fay. The stronger the acid, the more power it has to react, the more H^+ it will give. The stronger acid will react MORE than the weaker acid so the equilibrium

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will lie on the weaker acid side.

Chapter 15 Review Acid Base Equilibria

Chapter 14 Review Acids Bases Mixed
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CHAPTER 14 REVIEW Acids and Bases
SECTION 3 SHORT ANSWER Answer the

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following questions in the space provided. acids are sour (do not taste). When linked together, the amino acids form proteins, which make up many structures in the body. Any percent may be changed to an equivalent fraction by dropping the percent symbol and writing the number ...

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